Roll No. | | | | | | | | | | | | | Total No. of Pages : 02

Total No. of Questions : 10

B.Pharma (2017 & Onwards) (Sem.–1)

PHARMACEUTICAL ANALYSIS-I

Subject Code : BP-102T

Paper ID : [74645]

Time : 3 Hrs. Max. Marks : 80

INSTRUCTIONS TO CANDIDATES :

1. **SECTION-A is COMPULSORY consisting of FIFTEEN questions carrying TWO marks each.**
2. **SECTION-B contains FIVE questions carrying FIVE marks each and students have to attempt any FOUR questions.**
3. **SECTION-C contains FOUR questions carrying TEN marks each and students have to attempt any THREE questions.**

SECTION-A

1. Explain briefly :

a. Reduction Potential.

b. Primary standards.

c. Equivalent weight of potassium permanganate on acid and alkaline medium.

d. What do you understand by digestion of precipitates? Give its advantages.

e. Starch is added near to the end point of titration, why?

f. Name four different end point determination methods for precipitation titrations.

g. Werner Coordination number.

h. Masking agents.

i. Relation of pH to potential.

j. Chelation.

k. Cell constant.

1. Define and give units of Specific conductance.

m. Give one name of reference and standard electrode each.

n. Difference between Iodimetry and Iodometry titrations.

o. Name two metal ion indicators.

SECTION-B

2. Factors affecting stability of complexes.

3. Co-precipitation and Post-precipitation.

4. Derive the Handerson-Hasselbalch equation.

5. Taking a suitable example, explain the titration of weak bases by non-aqueous titrations.

6. Give a schematic diagram of assembly used in Arsenic limit test.

SECTION-C

7. Various types of electrochemical method of analysis, giving advantages and disadvantages.

8. Give the principle detailed procedure, reactions and use of each reagent used in Iron limit test.

9. Give a detailed account of sources and type of errors in pharmaceutical analysis.

10. Explain the constructions and working of dropping mercury electrode.